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S.4 CHEMISTRY BY MR. BOGERE.S.M

1. a) Define the term **MOLE**

b) Work out the number of moles in each of the following:

i) **4.5g** of hydrated copper(II)sulphate

ii) **480 cm³** of oxygen gas at room temperature and pressure

iii) **3.01 * 10²³** atoms of magnesium

iv) **200 cm³** of a **0.5M** sodium hydroxide solution

2. A certain gas is identified by its **rotten-egg smell** and its ability to form **misty**

Fumes in moist air

a) Name the gas

b) With the aid of a well labelled diagram, describe how a dry sample of this gas can be prepared in the laboratory.

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